

A quick plea...

Phases and Phase Changes - labels and definitions

G \longrightarrow L condensation

L \longrightarrow G boiling or evaporation
 evaporation - surface only
 boiling - throughout

L \longrightarrow S freezing

S \longrightarrow L melting

G \longrightarrow S deposition

S \longrightarrow G sublimation

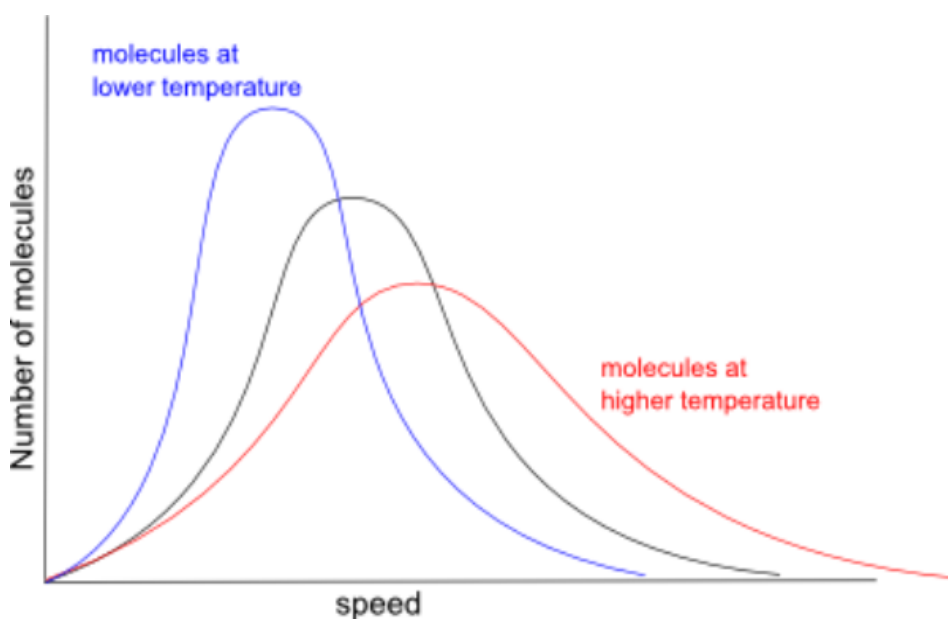
Heat v. Temperature

Heat (noun) - the total amount of kinetic energy that the particles have together

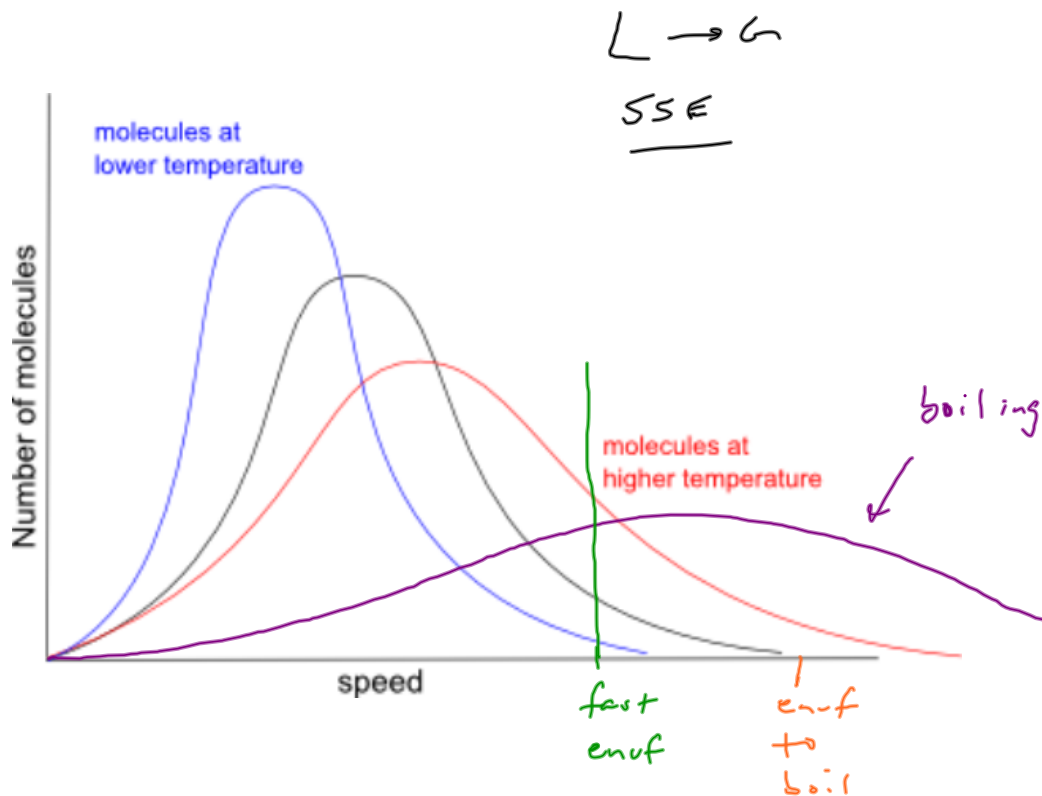
Temperature (adjective?) - a measure of the average kinetic energy

related b/c ... as heat is added to something, the total kinetic energy rises, forcing the average KE to also rise, and vice versa

Boltzmann Diagrams



Temperature and phase changes



Heat and phase changes

At what phase is water at 0°C?

solid or liquid or both

So what's the difference?

PE

