

States of Matter

solid - particles are locked in place relative to each other

liquid - move relative to each other but stay together due to attr.

gas - particles move independently

4 properties of gases (definitions, symbols and units)

Volume (V) - size of container - L

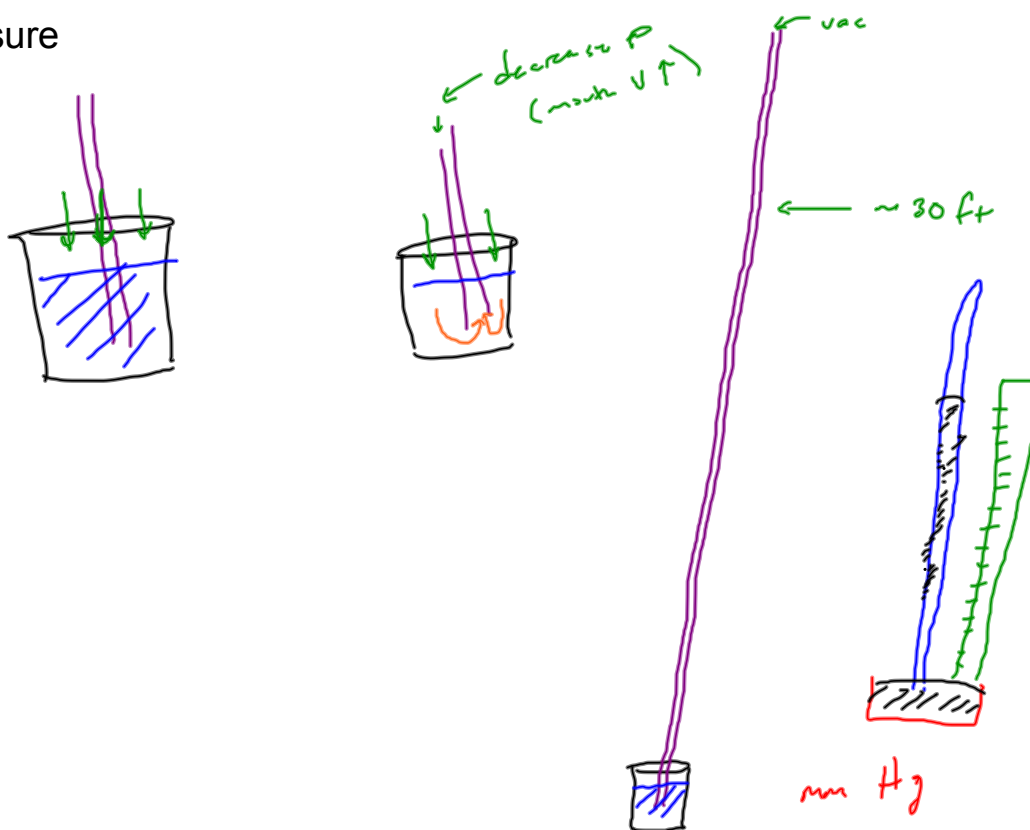
amount (n) - how much stuff - mols

Temperature (T) - measure of ave KE - K

Pressure (P) - force/area

↑
particles colliding
with surface

Pressure



$$1 \text{ mmHg} = 1 \text{ Torr}$$

$$760 \text{ mmHg} = 1 \text{ atm}$$

$$\text{Pascal (Pa)} = \frac{1 \text{ N}}{\text{m}^2}$$

$$101.325 \text{ kPa} = 1 \text{ atm}$$

Some reminders (vacuum and opposing forces)

vacuum - nothing

