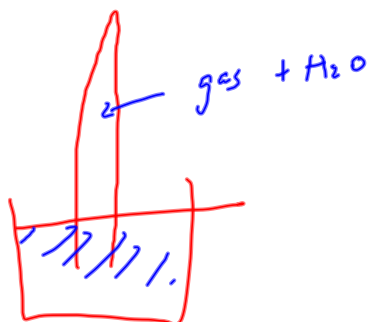
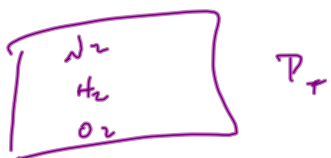
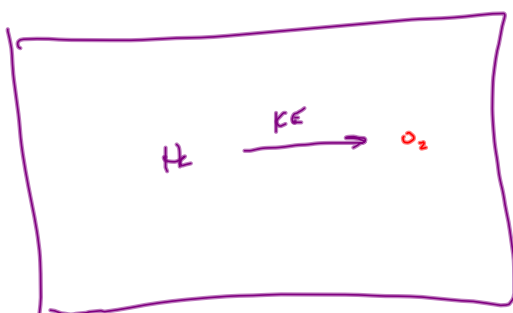


Dalton



$$P_{\text{atm}} = P_{\text{gas}} + P_{H_2O}$$



i	F
10.6L	?L
1.28atm	4.97atm
397K	308K

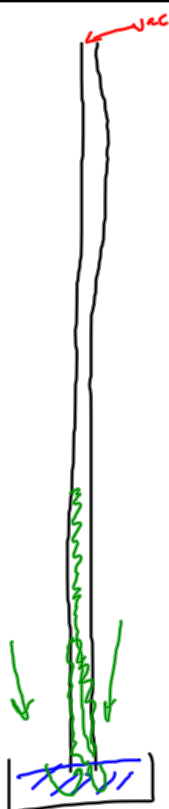
PTV ↓

T ↓ V ↓

$$10.6L \times \frac{1.28atm}{4.97atm} \times \frac{308K}{397K} =$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$\frac{P_1 V_1 T_2}{P_2 T_1} = V_2$$

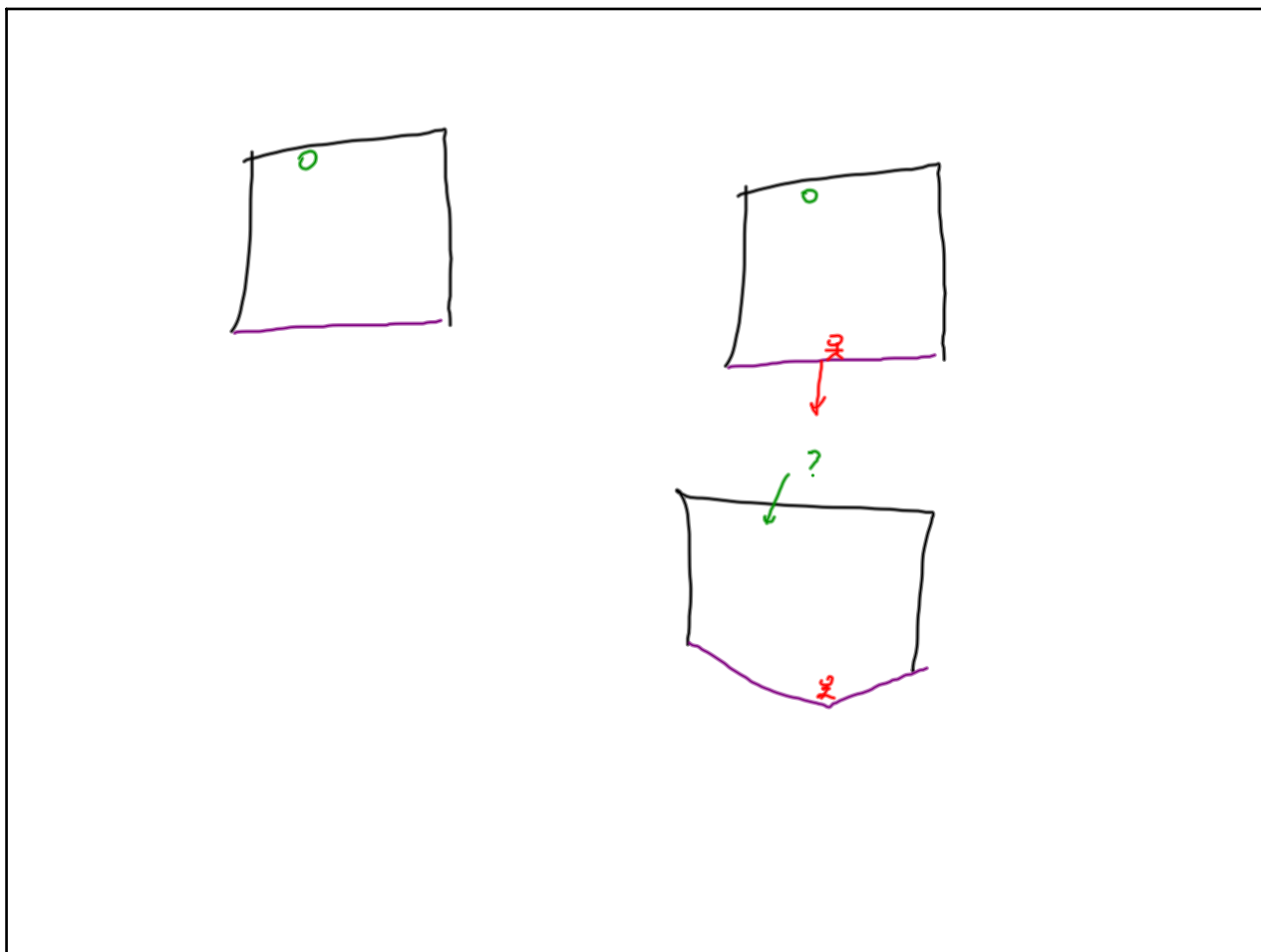


$$PV = nRT$$

$$PV = k$$

$$V = kT$$

$$V = \frac{nRT}{P}$$



$$\left[P + a \left(\frac{n}{V} \right)^2 \right] [V - nb] = nRT$$

\uparrow \uparrow
 P_i $V_i = nRT$

CO₂ 3 min 47.00 sec = 227.00 sec
 N₂ ?

$$\frac{\sqrt{\mu_{m1}}}{\sqrt{\mu_{m2}}} = \left(\frac{d_2}{d_1}\right) = \frac{\cancel{d_2} \cdot t_1}{t_2 \cdot \cancel{d_1}}$$

$$\frac{\sqrt{\mu_{CO_2}}}{\sqrt{\mu_{N_2}}} = \frac{t_{CO_2}}{t_{N_2}} \quad t_{N_2} = t_{CO_2} \frac{\sqrt{\mu_{N_2}}}{\sqrt{\mu_{CO_2}}}$$

$$(227.00 \text{ sec}) \left(\frac{\sqrt{28.0134}}{\sqrt{44.0095}} \right) = t_{N_2}$$

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