

Phases and Phase Changes - labels and definitions

G → L condensation

L → G boiling or evaporation
 evaporation - surface only
 boiling - throughout

L → S freezing

S → L melting

G → S deposition

S → G sublimation

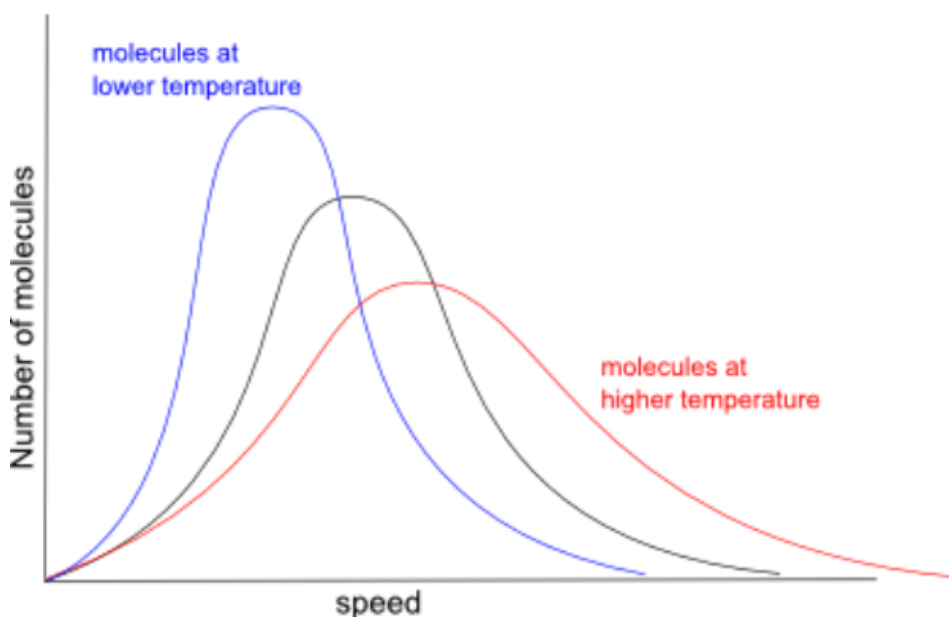
Heat v. Temperature

Heat (noun) - the total amount of kinetic energy that
 the particles have together

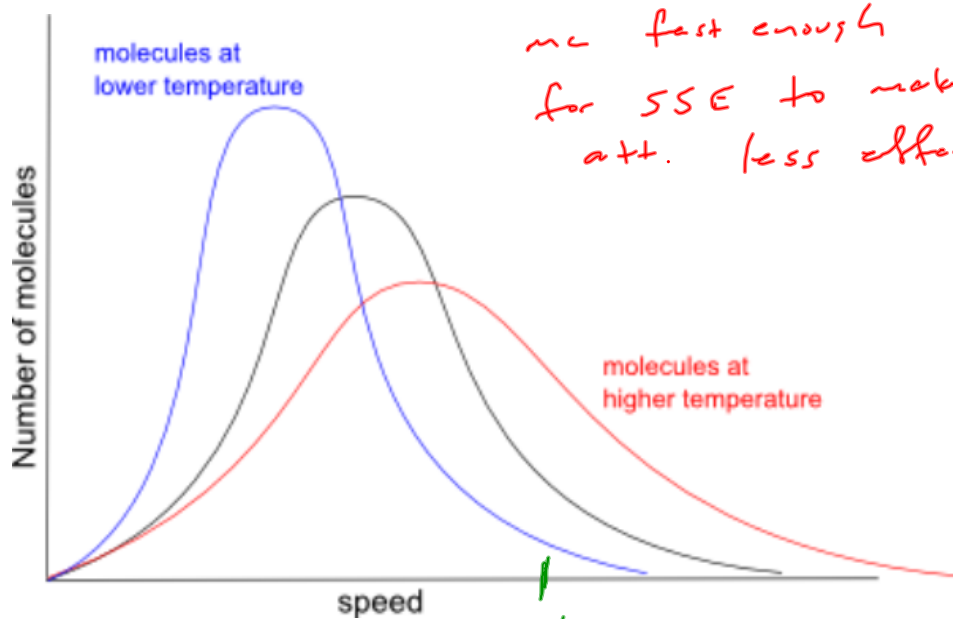
Temperature (adjective?) - a measure of the
 average kinetic energy

related b/c ... as heat is added to something, the
 total kinetic energy rises, forcing
 the average KE to also rise, and
 vice versa

Boltzmann Diagrams



Temperature and phase changes



Heat and phase changes

At what phase is water at 0°C ?

solid (ice) or liquid (water)

So what's the difference?

Potential NRG

