

pre-lab for Friday

goal

basic procedure

filtration

finding the lab

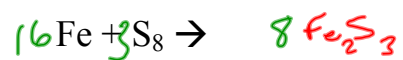


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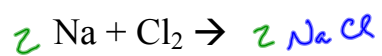
Synthesis

→ one thing

Charge of products...



Bold choice



Sep 4-6:13 PM

Decomposition

Change of products...



Sep 4-6:14 PM

Special Case Decomposition

Change of products...

no Δ 

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Special Case Decomposition

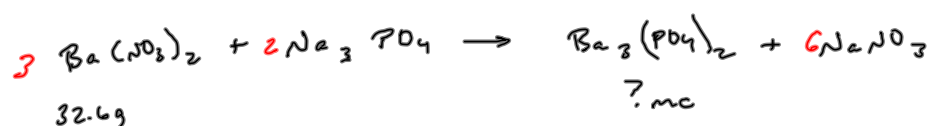
Charge of products...



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Stoichiometry

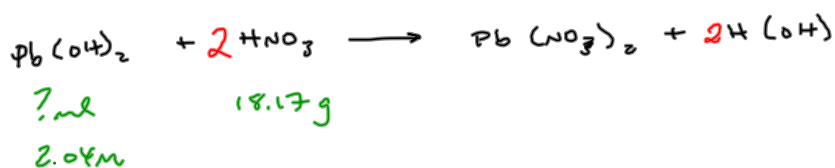
When 32.6 g of barium nitrate are reacted with sodium phosphate, how many molecules of the insoluble product are produced?



$$32.6 \text{ g Ba}(\text{NO}_3)_2 \times \frac{1 \text{ mol}}{261.337 \text{ g}} \times \frac{1 \text{ Ba}_3(\text{PO}_4)_2}{3 \text{ Ba}(\text{NO}_3)_2} \times \frac{6.022 \times 10^{23} \text{ mc}}{1 \text{ mol}} = 2.50 \times 10^{22} \text{ mc Ba}_3(\text{PO}_4)_2$$

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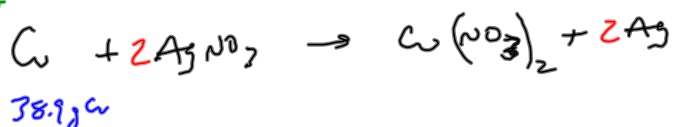
How many milliliters of a 2.04 M solution of lead II hydroxide are required to react completely with 18.17 g of nitric acid?



$$18.17 \text{ g HNO}_3 \times \frac{1 \text{ mol}}{63.0128 \text{ g}} \times \frac{1 \text{ Pb(OH)}_2}{2 \text{ HNO}_3} \times \frac{1000 \text{ ml}}{2.04 \text{ mol}} = 70.7 \text{ ml}$$

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If a 38.9 g piece of copper is dropped into a solution of silver nitrate and 38.9 g of silver are produced, what is the percent yield?



$$38.9 \text{ g Cu} \times \text{---} \times \text{---} \times \text{---} = \text{g Ag}$$

$$\% \text{ yield} = \frac{\text{got}}{\text{should get}} \times 100 = \frac{38.9 \text{ g Ag}}{\text{g Ag}} \times 100$$

mark answer
same units
= got

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